



K **Revision**



**HELPS YOU KNOW WHAT
TO STUDY NOW!!**

**Print it, do it anywhere
, and mark it done**

**Get rid of 100's of
resources**

Start calming parents

**Make studying
Fun again.!!!**

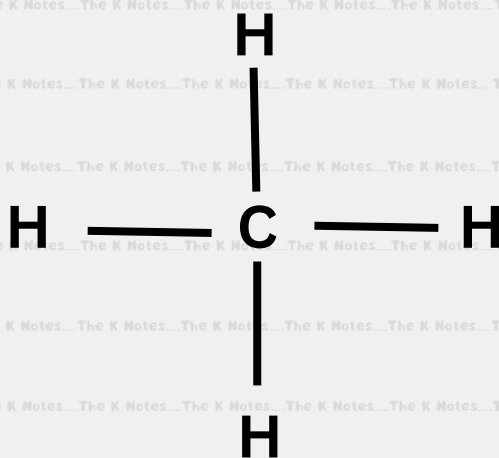
**Prevent Study
evaporation...**

**30 minutes weekday
60 minutes weekends
Find a time and Shine**

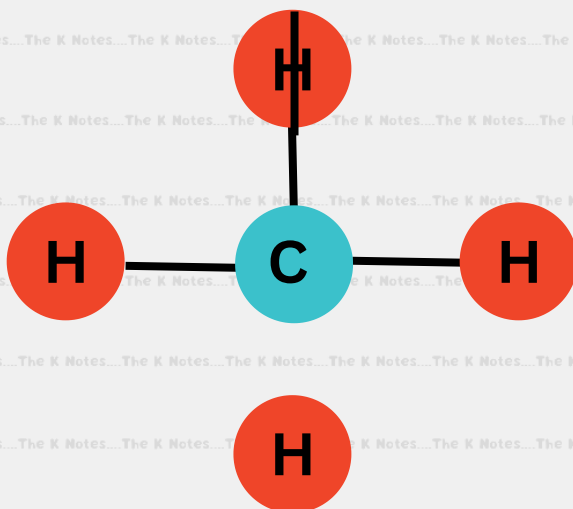
Week 1 Day 1

Displayed formulae

Shows how the atoms are bonded together as a diagram. For example Methane CH_4



Can also be shown as ball and stick diagram but not required to draw in our syllabus.



Week 1 Day 1

concentration in milligrams/litre

EXTRA INFORMATION FOR A PAST PAPER

sometimes the concentration of a certain Solute is measured in how many milligrams of the solute in a litre of the solution.

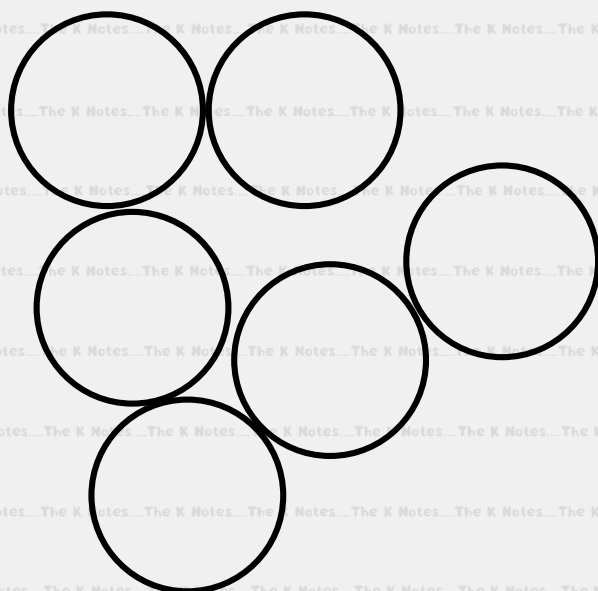
For example 20 mg/L means 20 mg of solute is dissolved in 1 litre of the solution.

ANALYSIS AT THE SOURCE			
TDS			47.6 mg/L
Total Hardness	40.0 mg/L	pH	6.8
Sodium			0.33 mg/L
Potassium	0.35 mg/L	Sulphate	1.9 mg/L
Calcium	9.6 mg/L	Fluoride	< 0.10 mg/L
Magnesium	4.0 mg/L	Chloride	0.15 mg/L
Nitrate	2.3 mg/L	Bicarbonate	49.0 mg/L
Added with Carbon Dioxide			
Free Carbon Dioxide (content at bottling)		min 5.6 g/L max 6.4 g/L	
تحليل في المصدر			
المواد الصلبة الذائبة الكلية		٤٧,٦ ملغ/لتر	
العسر الكلي		٤٠,٠ ملغ/لتر	
رقم الهيدروجيني		٦,٨	
صوديوم		٠,٣٣ ملغ/لتر	
بوتاسيوم		٠,٣٥ ملغ/لتر	
كبريت		١,٩ ملغ/لتر	
كالكسيوم		٩,٦ ملغ/لتر	
فلورايد		> ٠,١٠ ملغ/لتر	
مغنيزيوم		٤,٠ ملغ/لتر	
كلورايد		٠,١٥ ملغ/لتر	
نترات		٢,٣ ملغ/لتر	
بيكربونات		٤٩,٠ ملغ/لتر	
مضاف مع ثاني أكسيد الكربون			
ثاني أكسيد الكربون (أثناء التعبئة) ٥,٦ غرام/لتر كحد أدنى و ٦,٤ غرام/لتر كحد أقصى			

Week 1 Day 1

Drawing a liquid in the shape of particles

Draw big circles touching each other with some spaces. you can use a coin to draw regular circles. do not shade the circles or color them. Do not draw a lot of circles , six might be enough



Week 1 Day 1

Formula of Ions

For some ions you do not need to memorize them, as you can see the charge on the ion directly from the periodic table.

- Group 1 elements** ----> **+1 ions.** ---->e.g. **Na⁺,Li⁺**
- Group 2 elements** ----> **+2 ions.** ---->e.g. **Mg⁺²,Ba⁺²**
- Group 3 elements** ----> **+3 ions.** ---->e.g. **Al⁺³,Ga⁺³**
- Group 5 elements** ----> **-3 ions.** ---->e.g. **N⁻³,P⁻³**
- Group 6 elements** ----> **-2 ions.** ---->e.g. **O⁻²,S⁻²**
- Group 7 elements** ----> **-1 ions.** ---->e.g. **Cl⁻,Br⁻**

For some ions you can not find the Ion Charge in the periodic table, but you can know the charge from the formula, for example **ZnCl₂**, this means the charge of Zn is **+2**

For Polyatomic ions (ions made from more than one type of atom) you need to memorize them

For example

Hydroxide ion	OH⁻	Charge is -1
Carbonate ion	CO₃⁻	Charge is -2
Sulphate ion	SO₄⁻	Charge is -2
Hydrogencarbonate ion	HCO₃⁻	Charge is -1
Nitrate ion	NO₃⁻	Charge is -1
Ammonium ion	NH₄⁺	Charge is -1

Understand which forms of ionic compounds conduct electricity and why ?

- Molten form
- Aqueous form

conduct



Molten and aqueous forms can conduct electricity Because the ions will be free to move and carry the charge.

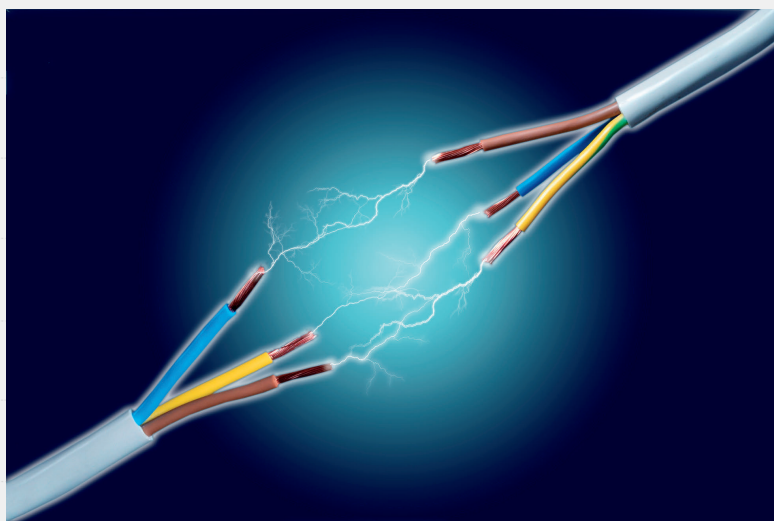
Understand which forms of ionic compounds does not conduct electricity and why ?

- Solid form

conduct



Solid form does not conduct electricity as the ions are held in fixed positions and cannot move.



Week 1 Day 2

order of reactivity of halogens

order of reactivity of halogens page

order of reactivity of group 1 page

order of reactivity of metals page

top most reactive bottom least reactive

compare reactivity of any 2 halogens (SKILL)

Uses of chlorine

**Chlorine Page , gas , solution ? colors , uses
formula**

reactions of chlorine with methane

Week 1 Day 6

Weekly Total Time Required To DO : 30 Minutes

Day 6 Content 30 mins